

Name: _____

Period: _____

Classifying Chemical Reactions

Questions To Ask:

Types of Reactions:

[]

Yes →

[]

No ↓

[]

Yes →

[]

No ↓

[]

Yes →

[]

No ↓

[]

Yes →

[]

No ↓

[]

Yes →

[]

Name: _____

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1. Displacement	A. To burn something.	1. Single Displacement Reaction	A. The metals in two compounds switch places.
2. Addition	B. To push something out of the way.	2. Addition Reaction	B. An element replaces one of the elements in a compound.
3. Combustion	C. To combine things.	3. Combustion Reaction	C. Two elements combine to form a compound.
4. Decomposition	D. To create something new from ingredients.	4. Double Displacement Reaction	D. A compound breaks up into its elements.
5. Synthesize	E. To break something apart.	5. Decomposition Reaction	E. A fuel burns in oxygen, creates heat, and usually produces CO ₂ .

Which type of reaction has only one reactant?

Which type of reaction has two compound reactants and two compound products?

Which type of reaction always has oxygen as a reactant?

Which type of reaction has only one product?

Which type of reaction has an element and a compound as reactants?

Write a reaction of Magnesium chloride combining with Lithium oxide to produce Magnesium oxide and Lithium Chloride. (Be sure that the ionic compounds are balanced.)

Directions: Complete the chart below.

Element	Number of Valence Electrons	# of electrons gained or lost to fill outer energy level	Charge (Oxidation Number)
Sodium			
Chlorine			
Beryllium			
Fluorine			
Lithium			
Oxygen			
Potassium			
Magnesium			
Phosphorous			
Aluminum			

Type of Reaction

Balance the reactions:

_____ Na₂S + _____ Zn(NO₃)₂ → _____ Na(NO₃) + _____ ZnS

_____ Li + _____ N₂ → _____ Li₃N

_____ KClO → _____ KCl + _____ O₂

_____ CH₄ + _____ O₂ → _____ H₂O + _____ CO₂

_____ Mg + _____ Ag(NO₃) → _____ Mg(NO₃)₂ + _____ Ag