












All sections marked with a  are considered essential concepts and must be completed to receive full credit on WS.

1. Oxidation #s	A. Attracted by a positive ion. 	6. Use the following symbols to answer the following.   
2. Zero	B. Tells you how many electrons will be gained or lost by an element.	
3. Negative ion	C. Net charge of a balanced ionic compound.	
4. Positive ion	D. When the number of electrons given equals the number taken.	
5. Balanced	E. Attracted by a negative ion.	
7. Give abbreviations with oxidation numbers and arrows 		8. Give number of electrons gained or lost 
Calcium	$\text{Ca}^{2+} \rightarrow$ Nitrogen	
Oxygen	Fluorine	
Sodium	Aluminum	
		Ca^{2+} <u>2 lost</u> Ca_3^{2+} <u>6 lost</u> F^{-} _____ F_3^{-} _____ Al^{3+} _____ Al_2^{3+} _____ O^{2-} _____ O_3^{2-} _____ Na^{1+} _____ Na_3^{1+} _____ N^{3-} _____ N_2^{3-} _____

For the following six examples, combine the two given atoms using electron arrows, then give the balanced ionic compound formula.

Combine Sodium and Oxygen 	Give the balanced ionic formula for Sodium Oxide.	Combine Beryllium and Fluorine 	Give the balanced ionic formula for Beryllium Fluoride.
Combine Magnesium and Sulfur 	Give the balanced ionic formula for Magnesium Sulfide	Combine Lithium and Phosphorus 	Give the balanced ionic formula: Compound name:
Combine Calcium and Nitrogen	Give the balanced ionic formula: Compound name:	Combine Aluminum and Oxygen	Give the balanced ionic formula for Compound name:
Write the balanced ionic compounds for the following: Be^{2+} and O^{2-} : _____ Na^{1+} and S^{2-} : _____ Li^{1+} and N^{3-} : _____ Ca^{2+} and N^{3-} : _____ Al^{3+} and Cl^{1-} : _____ K^{1+} and $(\text{SO}_4)^{2-}$: _____		Write the balanced ionic formulas for the following: Lithium and Oxygen: _____ Magnesium and Iodine: _____ Calcium and Sulfur: _____ Aluminum and Oxygen: _____	