

1. Oxidation #s B	<input checked="" type="checkbox"/> A. Attracted by a positive ion.	6. Use the following symbols to answer the following.
2. Zero C	<input checked="" type="checkbox"/> B. Tells you how many electrons will be gained or lost by an element.	$\rightarrow \quad \rightarrow \quad \rightarrow$
3. Negative ion A	<input checked="" type="checkbox"/> C. Net charge of a balanced ionic compound.	A. An electron being lost: \rightarrow
4. Positive ion E	<input checked="" type="checkbox"/> D. When the number of electrons given equals the number taken.	B. An ionic bond: $\rightarrow \text{---}$
5. Balanced D	<input checked="" type="checkbox"/> E. Attracted by a negative ion.	C. An electron being gained: \leftarrow
		D. Used for a metal: \rightarrow
		E. Used for a nonmetal: \leftarrow

7. Give abbreviations with oxidation numbers and arrows	8. Give number of electrons gained or lost
Calcium $\text{Ca}^{2+} \rightarrow \rightarrow$	Ca^{2+} <u>2 lost</u>
Nitrogen $\text{N}^{3-} \leftarrow \leftarrow \leftarrow$	Ca_3^{2+} <u>6 lost</u>
Oxygen $\text{O}^{2-} \leftarrow \leftarrow$	F^{1-} <u>1 G</u>
Fluorine $\text{F}^{1-} \leftarrow$	F_3^{1-} <u>3 G</u>
Sodium $\text{Na}^{1+} \rightarrow$	Al^{3+} <u>3 L</u>
Aluminum $\text{Al}^{3+} \rightarrow \rightarrow \rightarrow$	Al_3^{3+} <u>6 L</u>
	O^{2-} <u>2 G</u>
	O_3^{2-} <u>6 G</u>
	Na^{1+} <u>1 L</u>
	Na_3^{1+} <u>3 L</u>
	N^{3-} <u>3 G</u>
	N_2^{3-} <u>6 G</u>

For the following six examples, combine the two given atoms using electron arrows, then give the balanced ionic compound formula.

Combine Sodium and Oxygen $\text{Na} \rightarrow \rightarrow \text{O}$ $\text{Na} \rightarrow$	Give the balanced ionic formula for Sodium Oxide. Na_2O	Combine Beryllium and Fluorine $\text{Be} \rightarrow \rightarrow \text{F}$ $\rightarrow \text{F}$	Give the balanced ionic formula for Beryllium Fluoride. BeF_2
Combine Magnesium and Sulfur $\text{Mg} \rightarrow \rightarrow \text{S}$ $\text{Mg} \rightarrow$	Give the balanced ionic formula for Magnesium Sulfide. MgS	Combine Lithium and Phosphorus $\text{Li} \rightarrow \rightarrow \text{P}$ $\text{Li} \rightarrow$ $\text{Li} \rightarrow$	Give the balanced ionic formula: Li_3P Compound name: Lithium Phosphide
Combine Calcium and Nitrogen $\text{Ca} \rightarrow \rightarrow \text{N}$ $\text{Ca} \rightarrow \rightarrow \text{N}$ $\text{Ca} \rightarrow$	Give the balanced ionic formula: Ca_3N_2 Compound name: Calcium Nitride	Combine Aluminum and Oxygen $\text{Al} \rightarrow \rightarrow \text{O}$ $\text{Al} \rightarrow \rightarrow \text{O}$ $\rightarrow \text{O}$	Give the balanced ionic formula for Beryllium Fluoride. Al_2O_3 Compound name:
Write the balanced ionic compounds for the following: Be ²⁺ and O ²⁻ : <u>BeO</u> Na ¹⁺ and S ²⁻ : <u>Na₂S</u> Li ¹⁺ and N ³⁻ : <u>Li₃N</u> Ca ²⁺ and N ³⁻ : <u>Ca₃N₂</u> Al ³⁺ and Cl ¹⁻ : <u>AlCl₃</u> K ¹⁺ and (SO ₄) ²⁻ : <u>K₂(SO₄)</u>		Write the balanced ionic formulas for the following: Lithium and Oxygen: <u>Li₂O</u> Magnesium and Iodine: <u>MgI₂</u> Calcium and Sulfur: _____ Aluminum and Oxygen: _____	