CHEMICAL BONDS

Unit 2 Retest

- Only for those who scored less than a C

GUIDED READING READ PAGES 327-333 AND ANSWER THE FOLLOWING QUESTIONS

1. What is the octet rule and why is it important to bonding?
2. What are the exceptions to the octet rule and why are they exceptions?
3. Would it be easier for Lithium to gain seven electrons to fill the second shell- or to lose one?
4. Why is the Lewis dot structure helpful for us to predict how an atom will bond?
5. Why does Sodium tend to form bonds that allow it to give up its single valence electron?
6. How are Ionic bonds different than Covalent bonds?
7. Why do Sodium atoms and Chlorine atoms become attracted to each other?
8. Why can't microorganisms digest plastic?

Electron Dot Structures Poster

Draw the Elements Below with their Lewis Dot Structures and label the Groups and Periods

	1	2	3	4	5	6	7	8
1	H							He
2	Li	Be	B	C	N	O	F	Ne
3	Na	Mg	Al	Si	P	S	Cl	Ar