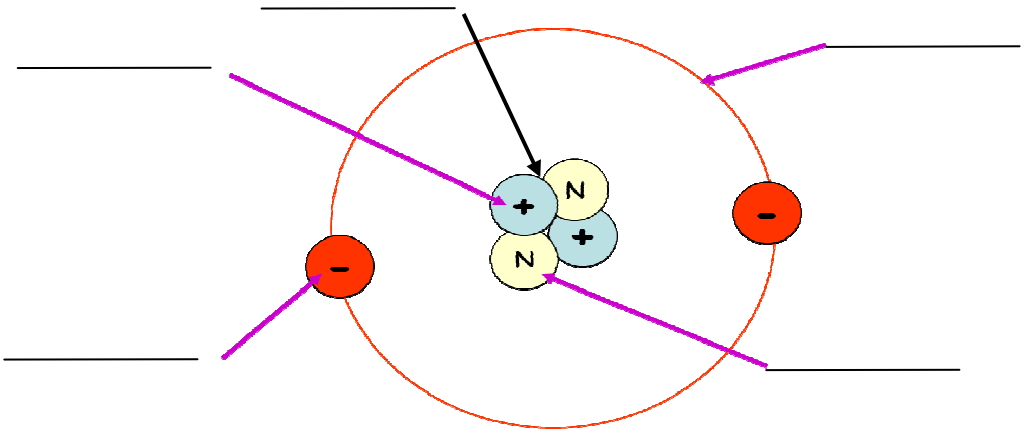
Name_____
Period & Group_____

ATOMIC BASICS

Subatomic Particles

Your Definintion

Particle	Charge	Location	Mass	Tells the



2

He

Helium

4

What does this tell you?

What does this tell you?

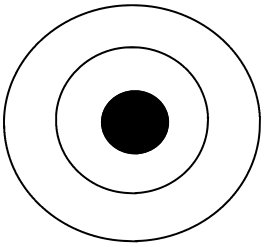
How to find # of Neutrons

_____

-

_____

_____# of Neutrons



7

N

Nitrogen

14

P_____

N_____

E_____

Electrons are found-

Max # of 1st Energy Level _____

Max # of 2nd Energy Level _____

Max # of 3rd Energy Level _____

Number of Electrons =

For the given Element find the Following:

_____

Atomic #

Chemical Symbol

Bohr Diagram

Atomic Mass

P_____

N_____

E_____

For the given Element find the Following:

_____

Atomic #

Chemical Symbol

Bohr Diagram

Atomic Mass

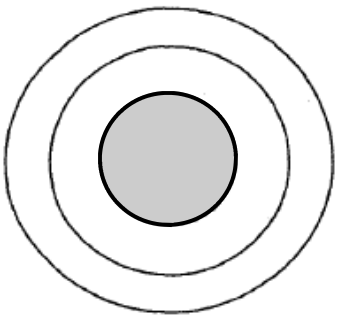
P_____

N_____

E_____

Atomic Basics Name _____

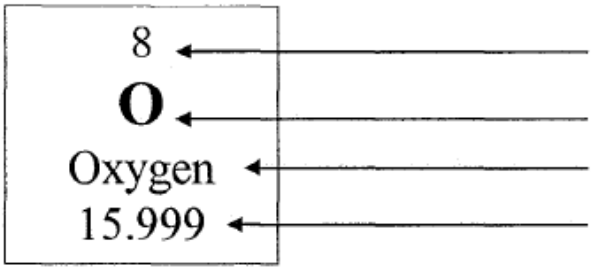
Part A: Atomic Structure



1. Draw five protons in the nucleus of the atom. Label them with their charge.
2. Draw six neutrons in the nucleus of the atom.
3. Draw two electrons in the first energy level and label them with their charge.
4. Draw three electrons in the second energy level and label them with their charge.
5. What element is represented by the diagram? _____

Part B: Atomic Calculations

6. Label the information provided in the periodic table.



7. What does the atomic number represent?

_____ or _____

8. What does the atomic mass represent?

_____ + _____

9. How would you figure the number of protons or electrons in an atom?

10. How would you figure the number of neutrons in an atom?

11. Use your knowledge of atomic calculations to complete the chart.

Element	Atomic Number	Atomic Mass	Protons	Neutrons	Electrons
Li	3	7			
P	15	31			
Cl		35	17		
Ni	28			31	
K		39			19
Ag	47			61	
H		1	1		
Si				14	14
W			74	110	
Ne				10	10